

## The Crystal Structure of *p*-Chlorobenzene *anti*-Diazoimidoglyoxynitrile

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The crystal structure of *p*-benzene-*anti*-diazoimidoglyoxynitrile,  $\text{ClC}_6\text{H}_4\cdot\text{N}_2\cdot\text{C}(\text{NH})\text{CN}$ , has been determined by three-dimensional X-ray methods. The crystals are monoclinic, space group  $P2_1/c$ , and the unit cell, containing four molecules, has the dimensions:

$$a = 19.223 \text{ \AA}; b = 5.943 \text{ \AA}; c = 7.523 \text{ \AA}; \beta = 91.47^\circ$$

The intensity data were collected using integrating Weissenberg techniques. 1273 independent reflections were used for the structure determination. Least-squares refinement procedures resulted in a conventional *R* factor of 9 %.

The molecule is nearly planar; the diazo group is of the *anti* type with bond angles at the diazo nitrogen atoms of 113 and 111°. Hydrogen bonds are present between imide groups of neighbouring molecules.

A diazocyanide is formed when a solution of KCN is mixed with an acid solution of a diazonium salt. If a neutral solution of the diazonium salt is added to a concentrated solution of potassium cyanide another compound is formed, a compound first prepared by Gabriel<sup>1</sup> in 1879 and having the composition  $\text{Ar}\cdot\text{N}_2\cdot\text{CN}_2\cdot\text{HCN}$ . It was regarded as a "double cyanide", but Hantzsch and Schulze,<sup>2</sup> resp. Hantzsch and Danziger<sup>3</sup> claimed that the compound probably was an imidocyanide,  $\text{Ar}\cdot\text{N}_2\cdot\text{C}(\text{NH})\text{CN}$ , of the diazocarboxylic acid because it may be reduced to a derivative of phenylhydrazine. Even in recent publications, however, the compound has been described as a true diazonium salt<sup>4</sup> and thus as the third of the isomers in the diazocyanide series. The present report deals with the crystal structure of the *p*-chlorobenzene diazo double cyanide, or, as it actually turned out, the *p*-chlorobenzene-*anti*-diazoimidonitrile.

### EXPERIMENTAL

The compound was prepared by adding an aqueous solution of *p*-chlorobenzene-diazonium chloride to a concentrated solution of potassium cyanide in a water-alcohol mixture at temperatures below  $-20^\circ\text{C}$ . After stirring for about half an hour the com-

found could be collected by filtration, and it was then purified by recrystallisation from absolute alcohol. (Found: C 50.11; H 2.65; N 29.51; Cl 17.99. Calc. for  $\text{C}_6\text{H}_4\text{N}_2\text{Cl}$ : C 49.88; H 2.62; N 29.09; Cl 18.41).

The unit cell dimensions were determined from Guinier films taken at room temperatures with  $\text{CuK}\alpha_1$  radiation ( $\lambda = 1.54050 \text{ \AA}$ ) and using powdered potassium chloride as a standard of calibration. The position of 13 well defined lines were used in a least-squares computation of the cell constants.

The crystals are of a brownish red colour and of prismatic form. The two specimens used in the X-ray experiments had the dimensions  $0.08 \times 0.12 \times 0.20 \text{ mm}^3$  and  $0.09 \times 0.08 \times 0.16 \text{ mm}^3$ , respectively. They were mounted for rotation about the direction of the largest extensions, which turned out to be parallel to the  $b$  and  $c$  axes. The intensity data were collected using multifilm integrating equi-inclination Weissenberg techniques with (Ni-filtered)  $\text{CuK}\alpha$ -radiation. The  $(hk0)$ -zone and the layers  $(h0l)$  to  $(h4l)$  were recorded and the intensities of all but the weakest reflections were measured photometrically. The weakest reflections were estimated visually using a calibrated scale. Of 1273 independent reflections which might be expected to appear on these films 816 were actually observed. The Lorentz and polarization factors were applied to the intensity data, but no absorption correction was performed.

The full-matrix least-squares program employed in the parameter refinement procedure was a modified version of the program written by Gantzel, Sparks and Trueblood (IUCr World List No. 384). The program minimizes the function  $\sum w^2(F_{\text{obs}} - G \cdot F_{\text{calc}})^2$ ; the weight applied to the structure factors was constant for  $|F_{\text{obs}}|$  less than 6 and proportional to  $|F_{\text{obs}}|^{-1/2}$  for larger structure factor values. Non-observed reflections were included in the calculations with a structure factor corresponding to the most probable value<sup>5</sup> and with the weight one third of that given to the observed reflections.

The atomic form factors used in the calculations were those derived by Hanson *et al.*<sup>6</sup>

#### CRYSTAL DATA

*p*-Chlorobenzene-*anti*-diazoidimido glyoxynitrile  $\text{Cl}\cdot\text{C}_6\text{H}_4\text{N}_2\text{C}(\text{NH})\text{CN}$ : Monoclinic, unit cell dimensions:

$$a = 19.223 (.006) \text{ \AA}; b = 5.943 (.002) \text{ \AA}; c = 7.523 (.001) \text{ \AA}; \beta = 91.47(.03)^\circ$$

Figures in parenthesis are standard deviations.

$$V = 859.2 \text{ \AA}^3, M = 192.62, F(000) = 392, Z = 4.$$

The density obtained by flotation is  $1.48 \text{ g}\cdot\text{cm}^{-3}$ , the calculated density is  $1.489 \text{ g}\cdot\text{cm}^{-3}$ .

$$\begin{aligned} \text{Absent reflections: } h0l & \text{ when } l = 2n + 1 \\ & 0k0 \text{ when } k = 2n + 1 \end{aligned}$$

$$\text{Space group: } P2_1/c$$

#### STRUCTURE DETERMINATION

The structure determination was started by taking advantage of the known approximate geometry of the *p*-chlorophenyl part of the molecule. The orientation and position of this group could easily be ascertained from the sharpened Patterson projections along the  $b$  and  $c$  axes. Subsequent Fourier syntheses yielded valuable information regarding the structure of the rest of the molecule. After a couple of Fourier refinements in the two projections a three-dimensional least-squares refinement was carried out.

After four cycles with variation of both positional and (isotropic) thermal parameters of the 13 non-hydrogen atoms, the conventional  $R$  index was 0.16.

The individual layer line data were rescaled, three additional cycles including also the benzene hydrogen atoms lowered  $R$  to 0.15. A new series of five least-squares cycles in which all non-hydrogen atoms were assigned thermal parameters, led to an  $R$  value of 0.10.

The position of the remaining hydrogen atom belonging to the imide group was determined by inspection of a three-dimensional Fourier difference map. A final least-squares refinement of all parameters except the thermal parameters of the hydrogen atoms ( $B=3.5 \text{ \AA}^2$ ) resulted in an  $R$  value of 0.09. 1273 structure factors were included in these calculations, giving an over-determination ratio of 9.6.

Table 1. Final positional parameters and their standard deviations (in parenthesis). The values have been multiplied by  $10^4$ .

	$x$	$y$	$z$
Cl	516 (1)	1948 (4)	1748 (3)
N(1)	2958 (3)	7943 (13)	1280 (10)
N(2)	3519 (3)	7248 (15)	1999 (10)
N(3)	3857 (3)	2485 (16)	0011 (11)
N(4)	4651 (3)	8280 (14)	2514 (10)
C(1)	2388 (3)	6378 (14)	1464 (9)
C(2)	2448 (3)	4283 (14)	2246 (10)
C(3)	1871 (3)	2951 (14)	2369 (10)
C(4)	1236 (3)	3703 (13)	1669 (10)
C(5)	1168 (3)	5787 (15)	0893 (10)
C(6)	1752 (3)	7164 (14)	0762 (10)
C(7)	4074 (3)	8845 (14)	1776 (11)
C(8)	3939 (3)	0856 (18)	0777 (13)
H(1)	2900 (47)	3540 (176)	2790 (125)
H(2)	1900 (49)	1440 (182)	2920 (122)
H(3)	0670 (44)	5890 (170)	0460 (112)
H(4)	1760 (55)	8610 (202)	0420 (151)
H(5)	4950 (57)	9520 (194)	2470 (140)

Table 2. Final thermal parameters for non-hydrogen atoms. The temperature factor is of the form  $\exp -(B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)$ . The values have been multiplied by  $10^5$ .

	$B_{11}$	$B_{22}$	$B_{33}$	$B_{12}$	$B_{13}$	$B_{23}$
Cl	213	4045	2967	-459	-109	648
N(1)	191	3204	2332	21	-136	535
N(2)	205	3468	2283	180	-225	119
N(3)	301	3450	3244	-81	-306	-325
N(4)	222	4824	3186	255	-294	404
C(1)	191	2503	2092	-175	-70	-189
C(2)	172	2847	2445	-11	-184	-252
C(3)	218	3555	2270	288	-96	-817
C(4)	212	2186	2493	55	75	-485
C(5)	168	3365	2375	43	-61	-949
C(6)	209	2914	2344	222	-141	-184
C(7)	188	2984	2706	-129	-32	-975
C(8)	191	3078	3005	-18	-138	-1550

Table 3. Observed and calculated structure factors. The columns are  $h, k, l, 10 |F_{\text{obs}}|, 10 F_{\text{calc}}$ . Asterisks indicate non-observed reflections.

1 0 0	119	116			3 4 0	133	139	17 0 2	129	-137	-8 0 8	14	26	2 1 2	314	307
2 0 0	105	-96			4 4 0	12	-18	18 0 2	50	-57	-7 0 8	15	25	3 1 2	269	-275
3 0 0	43	41			5 4 0	35	56	19 0 2	26	-23	-6 0 8	15	13	4 1 2	149	135
4 0 0	336	-353			6 4 0	207	188	20 0 2	135	-119	-5 0 8	15	31	5 1 2	150	152
5 0 0	254	275			7 4 0	81	-85	21 0 2	148	-146	-4 0 8	103	72	6 1 2	273	176
6 0 0	320	-337			8 4 0	146	-125	22 0 2	67	-74	-3 0 8	82	-5	7 1 2	442	470
7 0 0	264	-278			9 4 0	140	-128	23 0 2	10	0	-2 0 8	64	-39	8 1 2	280	294
8 0 0	49	-36			10 4 0	13	-26	24 0 2	23	-43	-1 0 8	66	-48	9 1 2	12	23
9 0 0	487	-504			11 4 0	119	121	-22 0 4	26	-42	0 0 8	31	-28	10 1 2	270	282
10 0 0	188	-195			12 4 0	28	-44	-21 0 4	12	-11	1 0 8	22	-25	11 1 2	95	85
11 0 0	252	-257			13 4 0	42	-38	-20 0 4	13	-25	2 0 8	74	-60	12 1 2	231	221
12 0 0	312	-321			14 4 0	13	21	-19 0 4	90	79	3 0 8	21	-38	13 1 2	11	6
13 0 0	89	-85			15 4 0	84	-79	-18 0 4	71	-68	4 0 8	15	12	14 1 2	86	-79
14 0 0	342	-341			16 4 0	12	-19	-17 0 4	69	-66	5 0 8	69	-61	15 1 2	309	-327
15 0 0	73	59			17 4 0	52	-50	-16 0 4	105	103	6 0 8	15	-3	16 1 2	232	-230
16 0 0	159	156			18 4 0	10	14	-15 0 4	127	106	7 0 8	14	14	17 1 2	98	102
17 0 0	451	452			19 4 0	9	19	-14 0 4	120	112	8 0 8	13	-4	18 1 2	158	-152
18 0 0	108	99			20 4 0	22	20	-13 0 4	173	158	9 0 8	13	20	19 1 2	90	-82
19 0 0	60	-52			1 5 0	20	-27	-12 0 4	168	153	10 0 8	12	22	20 1 2	82	-72
20 0 0	171	167			2 5 0	139	121	-11 0 4	323	303	11 0 8	43	36	21 1 2	76	-72
21 0 0	85	74			3 5 0	45	53	-10 0 4	128	-110	12 0 8	26	36	22 1 2	11	-12
22 0 0	14	12			4 5 0	139	94	-9 0 4	140	-122	13 0 8	82	74	23 1 2	9	20
23 0 0	11	43			5 5 0	133	-111	-8 0 4	15	17	-24 1 1	16	26	-22 1 3	8	10
24 0 0	16	24			6 5 0	61	-69	-7 0 4	204	-192	-23 1 1	83	78	-22 1 3	30	-32
1 1 0	149	142			7 5 0	82	70	-6 0 4	130	-122	-22 1 1	27	36	-21 1 3	30	34
2 1 0	489	-544			8 5 0	20	-23	-5 0 4	89	-86	-21 1 1	13	24	-20 1 3	107	-99
3 1 0	495	-541			9 5 0	35	-39	-4 0 4	260	-255	-20 1 1	119	116	-19 1 3	103	-100
4 1 0	313	-315			10 5 0	15	-5	-3 0 4	279	-260	-19 1 1	15	-20	-18 1 3	26	-35
5 1 0	307	-306			11 5 0	13	0	-2 0 4	39	29	-18 1 1	27	-25	-17 1 3	15	2
6 1 0	690	-755			12 5 0	87	75	-1 0 4	442	-414	-17 1 1	162	-65	-16 1 3	87	81
7 1 0	78	74			13 5 0	128	-95	0 0 4	37	-41	-16 1 1	22	-27	-15 1 3	39	-31
8 1 0	372	-378			14 5 0	78	-71	1 0 4	35	-43	-15 1 1	121	111	-14 1 3	15	1
9 1 0	9	-1			15 5 0	26	14	2 0 4	65	62	-14 1 1	65	-59	-13 1 3	44	44
10 1 0	10	6			16 5 0	33	-22	3 0 4	38	37	-13 1 1	14	-7	-12 1 3	14	-12
11 1 0	11	-13			17 5 0	28	-43	4 0 4	299	289	-12 1 1	148	-147	-11 1 3	24	19
12 1 0	208	206			1 6 0	59	50	5 0 4	472	432	-11 1 1	23	10	-10 1 3	83	-62
13 1 0	211	194			2 6 0	59	37	6 0 4	336	330	-10 1 1	107	-99	-9 1 3	154	147
14 1 0	258	247			3 6 0	13	-18	7 0 4	68	-56	-9 1 1	420	-433	-8 1 3	298	300
15 1 0	350	352			4 6 0	103	75	8 0 4	55	56	-8 1 1	10	-31	-7 1 3	56	-57
16 1 0	56	-55			5 6 0	54	-54	9 0 4	165	162	-7 1 1	263	268	-6 1 3	174	171
17 1 0	63	51			6 6 0	27	14	10 0 4	66	53	-6 1 1	95	93	-5 1 3	18	-18
18 1 0	200	189			7 6 0	11	-21	11 0 4	16	18	-5 1 1	339	374	-4 1 3	10	11
19 1 0	73	-70			8 6 0	126	-84	12 0 4	30	-39	-4 1 1	41	24	-3 1 3	294	-303
20 1 0	75	68			9 6 0	11	5	13 0 4	106	-101	-3 1 1	650	730	-2 1 3	373	-383
21 1 0	62	-62			10 6 0	79	-46	14 0 4	238	-228	-2 1 1	10	-192	-1 1 3	89	77
22 1 0	16	-14			11 6 0	10	-8	15 0 4	17	26	0 1 1	36	-28	0 1 3	9	8
23 1 0	87	-86			12 6 0	15	-19	16 0 4	158	-155	1 1 1	5	19	1 1 3	57	-60
24 1 0	62	-63			13 6 0	9	9	17 0 4	99	-94	2 1 1	5	-33	2 1 3	44	-44
0 2 0	905	-956			14 6 0	41	42	18 0 4	63	-60	3 1 1	169	-173	3 1 3	35	-28
1 2 0	82	-88			1 7 0	15	-6	19 0 4	99	-106	4 1 1	1	-39	4 1 3	27	13
2 2 0	67	-58			2 7 0	44	-23	20 0 4	18	-15	5 1 1	51	47	5 1 3	26	-35
3 2 0	239	-248			3 7 0	21	-8	21 0 4	50	58	6 1 1	177	-182	6 1 3	134	-120
4 2 0	59	56			4 7 0	42	-22	22 0 4	48	65	7 1 1	296	-291	7 1 3	104	84
5 2 0	146	-131			5 7 0	31	-32	23 0 4	9	21	8 1 1	338	-357	8 1 3	125	120
6 2 0	270	-256			6 7 0	47	-39	24 0 4	18	21	9 1 1	338	-357	9 1 3	228	225
7 2 0	252	234			7 7 0	20	-21	25 0 4	13	19	10 1 1	70	73	10 1 3	101	92
8 2 0	420	390			8 7 0	14	18	26 0 4	5	5	11 1 1	239	238	11 1 3	86	72
9 2 0	373	365			9 7 0	43	53	27 0 4	15	-10	12 1 1	121	-123	12 1 3	15	-22
10 2 0	188	175			10 7 0	2	-58	28 0 4	16	-10	13 1 1	140	142	13 1 3	100	-82
11 2 0	60	-50			11 7 0	15	15	29 0 4	16	-2	14 1 1	130	126	14 1 3	100	-82
12 2 0	232	229			12 7 0	106	-108	30 0 4	29	-31	15 1 1	99	91	15 1 3	39	-39
13 2 0	78	68			13 7 0	63	-59	31 0 4	226	-190	16 1 1	15	-2	16 1 3	15	-73
14 2 0	28	30			14 7 0	233	240	32 0 4	119	-119	17 1 1	111	-101	17 1 3	15	-33
15 2 0	19	42			15 7 0	278	-270	33 0 4	27	-24	18 1 1	92	85	18 1 3	50	-40
16 2 0	19	-21			16 7 0	293	-283	34 0 4	17	-25	19 1 1	71	62	19 1 3	45	-26
17 2 0	143	-153			17 7 0	293	-280	35 0 4	38	-28	20 1 1	40	39	20 1 3	13	25
18 2 0	48	-53			18 7 0	180	-172	36 0 4	17	16	21 1 1	14	7	21 1 3	46	-45
19 2 0	28	-35			19 7 0	48	51	37 0 4	17	-23	22 1 1	13	17	22 1 3	9	-27
20 2 0	65	-74			20 7 0	122	-117	38 0 4	17	-19	23 1 1	12	7	23 1 3	16	-20
21 2 0	66	-68			21 7 0	118	116	39 0 4	274	226	24 1 1	20	-32	24 1 3	25	34
22 2 0	57	-68			22 7 0	426	423	40 0 4	120	91	25 1 1	75	-20	25 1 3	59	62
23 2 0	63	-73			23 7 0	12	4	41 0 4	116	102	26 1 2	14	-14	26 1 3	62	-61
1 3 0	80	-90			24 7 0	443	450	42 0 4	74	71	27 1 2	72	-68	27 1 3	69	68
2 3 0	169	148			25 7 0	192	211	43 0 4	40	40	28 1 2	72	-81	28 1 3	139	135
3 3 0	287	272			26 7 0	206	222	44 0 4	100	90	29 1 2	70	-69	29 1 3	116	114
4 3 0	10	23			27 7 0	340	349	45 0 4	100	89	30 1 2	14	-24	30 1 3	97	90
5 3 0	215	181			28 7 0	170	-150	46 0 4	17	7	31 1 2	156	-155	31 1 3	32	-26
6 3 0	222	186			29 7 0	510	523	47 0 4	23	11	32 1 2	82	-79	32 1 3	16	12
7 3 0	11	1			30 7 0	163	154	48 0 4	248	-216	33 1 2	16	21	33 1 3	15	-29
8 3 0	119	96			31 7 0	1008	1021	49 0 4	86	-82	34 1 2	41	42	34 1 3	110	-105
9 3 0	21	-32			32 7 0	762	-731	50 0 4	17	-3	35 1 2	57	52	35 1 3	37	-26
10 3 0	65	53			1 0 2	1204	-1232	9 0 6	109	-112	36 1 2	178	182	36 1 3	150	-149
11 3 0	53	59			2 0 2	311	-296	10 0 6	75	-72	37 1 2	169	163	37 1 3	266	-266
12 3 0	70	-63			3 0 2	404	-407	11 0 6	86	-77	38 1 2	132	115	38 1 3	102	-89
13 3 0	31	-33			4 0 2	772	-769	12 0 6	90	-81	39 1 2	306	319	39 1 3	148	-139
14 3 0	121	-110			5 0 2	72	-59	13 0 6	38	-38	40 1 2	174	-166	40 1 3	19	18
15 3 0	119	-104			6 0 2	200	-214	14 0 6	128	123	41 1 2	304	312	41 1 3	140	-142
16 3																

Table 3. Continued.

6	1	4	48	-47	-6	1	7	48	36	19	2	1	129	-107	-9	2	4	19	44	-12	3	7	16	-3
7	1	4	93	85	-5	1	7	15	7	20	2	1	33	-45	-8	2	4	18	-56	-11	2	7	54	-49
8	1	4	294	-295	-4	1	7	15	14	21	2	1	96	-84	-7	2	4	110	107	-10	2	7	56	51
9	1	4	37	-33	-3	1	7	15	15	22	2	1	30	-21	-6	2	4	103	105	-5	2	7	58	53
10	1	4	116	-101	-2	1	7	15	11	-22	2	2	13	23	-5	2	4	31	-40	-8	2	7	19	31
11	1	4	166	-164	-1	1	7	15	21	-20	2	2	61	70	-4	2	4	332	360	-7	2	7	19	30
12	1	4	157	-130	0	1	7	15	1	-19	2	2	72	80	-3	2	4	256	257	-6	2	7	19	37
13	1	4	148	-140	0	1	7	15	1	-20	2	2	19	2	-2	2	4	59	-59	-5	2	7	19	37
14	1	4	192	-190	2	1	7	15	-20	-17	2	2	131	132	-1	2	4	185	178	-4	2	7	52	-69
15	1	4	50	37	3	1	7	23	-24	-17	2	2	50	-53	0	2	4	16	-45	-3	2	7	82	-79
16	1	4	158	198	-4	1	7	15	-1	-16	2	2	137	132	1	2	4	72	-71	-2	2	7	20	3
17	1	4	14	-3	5	1	7	56	-42	-15	2	2	99	93	2	2	4	144	-132	0	2	7	20	-30
18	1	4	52	32	6	1	7	15	-22	-14	2	2	56	66	4	2	4	35	-44	1	2	7	20	-22
19	1	4	75	68	7	1	7	15	4	-13	2	2	19	-8	-1	2	4	180	-193	2	2	7	20	-22
20	1	4	84	76	8	1	7	21	-33	-11	2	2	59	53	6	2	4	173	-172	3	2	7	20	24
21	1	4	21	23	10	1	7	14	-17	-10	2	2	50	-53	7	2	4	18	-20	4	2	7	19	-54
-21	1	5	0	-21	11	1	7	13	27	-9	2	2	16	19	8	2	4	79	-74	5	2	7	19	-57
-20	1	5	0	-21	12	1	7	12	5	-8	2	2	138	-109	9	2	4	86	-91	6	2	7	19	-45
-19	1	5	25	36	13	1	7	11	-23	-7	2	2	474	-454	10	2	4	72	-75	7	2	7	19	56
-17	1	5	13	28	14	1	7	10	20	-6	2	2	76	-57	11	2	4	71	-68	8	2	7	89	70
-15	1	5	13	28	15	1	7	9	12	-5	2	2	219	-201	12	2	4	20	51	9	2	7	17	11
-14	1	5	15	31	-13	1	8	57	55	-2	2	2	244	-222	13	2	4	124	132	10	2	7	17	47
-13	1	5	32	31	-12	1	8	53	59	-3	2	2	11	11	14	2	4	159	170	11	2	7	16	44
-12	1	5	16	-20	-11	1	8	10	-2	-2	2	2	99	101	15	2	4	19	-34	19	2	7	13	-39
-11	1	5	15	7	-10	1	8	11	5	-1	2	2	116	-100	16	2	4	57	73	13	2	7	12	-14
-10	1	5	16	20	-9	1	8	12	-2	0	2	2	244	227	-17	2	4	79	94	14	2	7	12	-14
-9	1	5	56	41	-7	1	8	13	-3	1	2	2	413	404	18	2	4	16	9	-9	2	8	14	-18
-8	1	5	107	-85	-6	1	8	13	-0	2	2	2	251	239	19	2	4	39	42	-8	2	8	15	-13
-7	1	5	33	-33	-5	1	8	36	-38	3	2	2	81	76	-19	2	5	13	23	-7	2	8	15	-6
-6	1	5	59	-58	-4	1	8	34	-23	4	2	2	293	290	-18	2	5	39	-43	-5	2	8	16	-29
-5	1	5	70	-67	-3	1	8	30	-19	5	2	2	284	263	-17	2	5	16	2	-4	2	8	17	-24
-4	1	5	99	-99	-2	1	8	61	-50	6	2	2	231	200	-16	2	5	18	14	-3	2	8	17	1
-3	1	5	14	-19	-1	1	8	14	-25	7	2	2	102	-96	-15	2	5	19	5	-2	2	8	17	12
-2	1	5	49	47	0	1	8	14	-21	8	2	2	16	14	-13	2	5	20	-52	0	2	8	17	20
-1	1	5	28	39	1	1	8	51	-35	10	2	2	253	-240	-12	2	5	20	1	1	2	8	17	35
0	1	5	13	9	2	1	8	14	-15	11	2	2	100	-76	-11	2	5	20	1	1	2	8	17	34
1	1	5	106	85	3	1	8	14	-5	12	2	2	168	-140	-10	2	5	157	-178	3	2	8	17	21
2	1	5	113	97	4	1	8	105	87	13	2	2	296	-279	-9	2	5	20	-12	4	2	8	16	19
3	1	5	14	-6	5	1	8	119	87	14	2	2	20	-17	-8	2	5	67	-56	5	2	8	16	34
4	1	5	14	27	6	1	8	47	-39	15	2	2	20	-17	-7	2	5	19	-39	6	2	8	15	4
5	1	5	14	11	7	1	8	40	26	16	2	2	207	-190	-6	2	5	19	11	7	2	8	47	-46
6	1	5	38	39	8	1	8	31	29	17	2	2	64	56	-5	2	5	49	79	8	2	8	100	-126
7	1	5	47	42	9	1	8	32	22	18	2	2	19	7	-4	2	5	277	296	-20	3	1	21	-59
8	1	5	15	3	10	1	8	32	34	19	2	2	18	7	-3	2	5	18	-15	-19	3	1	21	-59
9	1	5	59	-57	11	1	8	9	13	20	2	2	44	45	-3	2	5	18	-18	-18	3	1	22	15
10	1	5	15	-9	-8	1	9	7	0	21	2	2	15	30	-17	2	5	18	-21	-17	3	1	39	73
11	1	5	64	-63	-7	1	9	8	14	22	2	2	73	70	0	2	5	58	62	-16	3	1	24	53
12	1	5	15	-17	-6	1	9	9	-7	-22	2	3	21	-49	1	2	5	114	114	-15	3	1	60	28
13	1	5	64	-59	-5	1	9	9	-21	-20	2	3	14	-33	2	2	5	18	21	-14	3	1	40	54
14	1	5	15	3	-4	1	9	10	2	-20	2	3	16	11	3	2	5	18	0	-13	3	1	108	107
15	1	5	27	-10	-3	1	9	10	2	-19	2	3	17	22	-2	2	5	18	4	-12	3	1	268	252
16	1	5	36	-36	-2	1	9	10	4	-18	2	3	49	49	4	2	5	19	44	-6	3	1	29	28
17	1	5	41	34	-1	1	9	10	-5	-17	2	3	19	-3	5	2	5	19	-39	-11	3	1	83	78
18	1	5	26	24	0	1	9	10	-16	-16	2	3	20	41	6	2	5	19	-74	-10	3	1	69	83
19	1	5	21	29	1	1	9	10	2	-15	2	3	20	28	7	2	5	232	-252	-9	3	1	140	130
20	1	5	8	9	2	1	9	10	-11	-14	2	3	144	144	8	2	5	20	-39	-8	3	1	18	10
-19	1	6	26	41	3	1	9	10	-9	-13	2	3	48	59	-7	3	5	20	-16	-7	3	1	172	-177
-18	1	6	24	-25	4	1	9	9	-11	-12	2	3	41	48	10	2	5	97	-105	-6	3	1	29	28
-17	1	6	67	-53	5	1	9	9	4	-11	2	3	138	147	11	2	5	20	40	-5	3	1	284	-290
-16	1	6	76	-67	6	1	9	8	31	-10	2	3	38	38	12	2	5	20	28	-4	3	1	227	-239
-15	1	6	52	-52	7	1	9	7	-2	-9	2	3	115	-119	13	2	5	105	118	-3	3	1	365	-415
-14	1	6	46	-40	-8	1	9	11	-0	-8	2	3	43	-42	14	2	5	19	-34	-2	3	1	72	-79
-13	1	6	106	-92	-9	1	9	11	8	-7	2	3	144	-132	15	2	5	18	0	-1	3	1	10	-115
-12	1	6	15	2	-22	2	1	16	-8	-6	2	3	138	-131	16	2	5	17	49	0	3	1	55	63
-11	1	6	66	52	-20	2	1	33	-47	-5	2	3	276	-276	17	2	5	15	50	1	3	1	63	-58
-10	1	6	15	-29	-19	2	1	19	-22	-4	2	3	184	-177	18	2	5	13	28	2	3	1	179	-169
-9	1	6	106	83	-17	2	1	175	-66	-3	2	3	210	183	-17	2	6	13	6	3	3	1	86	88
-8	1	6	62	62	-18	2	1	136	-141	-2	2	3	85	-90	-16	2	6	14	-4	4	3	1	186	159
-7	1	6	56	53	-16	2	1	104	-92	-2	2	3	84	88	-15	2	6	16	9	5	3	1	425	410
-6	1	6	63	62	-15	2	1	64	-56	0	2	3	345	-327	-14	2	6	17	45	6	3	1	118	99
-5	1	6	88	75	-14	2	1	80	-70	1	2	3	13	33	-13	2	6	18	17	7	3	1	203	192
-4	1	6	15	-4	-13	2	1	34	37	2	2	3	119	107	-12	2	6	19	26	8	3	1	256	237
-3	1	6	159	141	-12	2	1	60	-52	3	2	3	179	181	-11									

Table 3. Continued.

1 3 2	211	225	16 3 3	22	20 m	0 3 5	23	-41 m	-13 h 1	30	-34 m	12 h 2	60	-58 m
2 3 2	51	38	17 3 3	80	77	1 3 5	23	-67 m	-12 h 1	29	69 m	-11 h 3	30	-21 m
3 3 2	52	-37	18 3 3	20	49 m	2 3 5	67	-103 m	-11 h 1	28	-14 m	-10 h 3	29	2 m
4 3 2	130	122	-17 3 4	20	2 m	3 3 5	23	-58 m	-10 h 1	217	-197 m	-9 h 3	29	64 m
5 3 2	104	-85	-16 3 4	21	-6 m	4 3 5	96	-113 m	-9 h 1	170	-142 m	-8 h 3	57	-70 m
6 3 2	227	-200	-15 3 4	22	-6 m	5 3 5	39	-57 m	-8 h 1	50	-42 m	-7 h 3	39	38 m
7 3 2	268	-236	-14 3 4	23	44 m	6 3 5	24	-19 m	-7 h 1	294	-264 m	-6 h 3	231	254 m
8 3 2	190	-169	-13 3 4	24	4 m	7 3 5	24	-66 m	-6 h 1	262	-230 m	-5 h 3	183	196 m
9 3 2	165	164	-12 3 4	24	-7 m	8 3 5	24	-42 m	-5 h 1	20	-29 m	-4 h 3	124	135 m
10 3 2	40	-42	-11 3 4	24	-17 m	9 3 5	130	162 m	-4 h 1	18	51 m	-3 h 3	25	-6 m
11 3 2	120	-118	-10 3 4	23	58 m	10 3 5	59	90 m	-3 h 1	105	128 m	-2 h 3	24	69 m
12 3 2	23	23 m	-9 3 4	228	254 m	11 3 5	39	69 m	-2 h 1	15	-13 m	-1 h 3	24	39 m
13 3 2	99	-91	-8 3 4	23	55 m	12 3 5	22	-18 m	-1 h 1	121	168 m	0 h 3	24	59 m
14 3 2	24	22 m	-7 3 4	22	-39 m	13 3 5	22	46 m	0 h 1	97	107 m	1 h 3	24	-57 m
15 3 2	24	25 m	-6 3 4	22	32 m	14 3 5	21	77 m	1 h 1	152	166 m	2 h 3	24	-35 m
16 3 2	23	29 m	-5 3 4	42	-50 m	15 3 5	19	14 m	2 h 1	127	130 m	3 h 3	221	-224 m
17 3 2	54	63	-4 3 4	21	-16 m	+13 3 6	19	-32 m	3 h 1	166	156 m	4 h 3	73	-74 m
18 3 2	21	46 m	-3 3 4	20	-58 m	-12 3 6	20	24 m	4 h 1	99	82 m	5 h 3	75	-82 m
19 3 2	109	86	-2 3 4	20	-30 m	-11 3 6	21	9 m	5 h 1	132	134 m	6 h 3	67	-67 m
-18 3 3	59	57	-1 3 4	20	-40 m	-10 3 6	22	60 m	6 h 1	22	33 m	7 h 3	40	-54 m
-17 3 3	22	18 m	0 3 4	34	-54 m	-9 3 6	23	-38 m	7 h 1	24	-53 m	8 h 3	110	-107 m
-16 3 3	23	1 m	1 3 4	20	6 m	-8 3 6	97	-116 m	8 h 1	25	-23 m	9 h 3	87	-83 m
-15 3 3	23	-6 m	2 3 4	129	-121 m	-7 3 6	23	31 m	9 h 1	96	-58 m	10 h 3	29	-8 m
-14 3 3	24	21 m	3 3 4	21	13 m	-6 3 6	83	-86 m	10 h 1	308	-259 m	11 h 3	161	185 m
-13 3 3	20	22	4 3 4	77	-71 m	-5 3 6	57	-84 m	11 h 1	307	-263 m	-7 h 3	60	70 m
-12 3 3	97	-120	5 3 4	215	-226 m	-4 3 6	24	17 m	12 h 1	29	33 m	-8 h 3	59	-75 m
-11 3 3	106	-124	6 3 4	22	20 m	-3 3 6	24	-10 m	13 h 1	30	-55 m	-6 h 3	72	-76 m
-10 3 3	60	-66	7 3 4	22	43 m	-2 3 6	106	90 m	+13 h 2	30	27 m	-5 h 3	58	48 m
-9 3 3	131	-122	8 3 4	193	206 m	-1 3 6	24	-16 m	-12 h 2	29	78 m	-4 h 3	134	-117 m
-8 3 3	229	-227	9 3 4	80	80 m	0 3 6	24	-14 m	-11 h 2	29	-29 m	-3 h 4	28	-28 m
-7 3 3	20	-18 m	10 3 4	89	-113 m	1 3 6	24	62 m	-10 h 2	28	-42 m	-2 h 4	28	27 m
-6 3 3	66	-75 m	11 3 4	151	166 m	2 3 6	24	-26 m	-9 h 2	27	-31 m	-1 h 4	27	31 m
-5 3 3	32	63	12 3 4	24	-10 m	3 3 6	24	32 m	-8 h 2	93	67 m	0 h 4	27	56 m
-4 3 3	18	-4 m	13 3 4	23	-45 m	4 3 6	24	-5 m	-7 h 2	223	222 m	1 h 4	28	-27 m
-3 3 3	195	204	14 3 4	77	88 m	5 3 6	24	51 m	-6 h 2	24	-32 m	2 h 4	28	-53 m
-2 3 3	169	183	15 3 4	22	-12 m	6 3 6	80	100 m	-5 h 2	23	-1 m	3 h 4	28	22 m
-1 3 3	98	102	16 3 4	49	65 m	7 3 6	23	21 m	-4 h 2	31	46 m	4 h 4	28	13 m
0 3 3	100	107	17 3 4	19	2 m	8 3 6	23	7 m	-3 h 2	82	-98 m	5 h 4	29	42 m
1 3 3	103	118	-15 3 5	20	-26 m	9 3 6	22	-39 m	-2 h 2	20	-42 m	6 h 4	29	7 m
2 3 3	146	148	-14 3 5	21	-23 m	10 3 6	22	11 m	0 h 2	92	-107 m	7 h 4	29	13 m
3 3 3	107	99	-13 3 5	22	-51 m	11 3 6	21	-40 m	1 h 2	19	65 m	8 h 4	30	43 m
4 3 3	150	138	-12 3 5	23	-14 m	12 3 6	94	-111 m	2 h 2	19	-7 m	-5 h 5	30	-80 m
5 3 3	140	-140	-11 3 5	23	39 m	-5 3 7	22	-31 m	3 h 2	20	17 m	-4 h 5	95	-80 m
6 3 3	58	-60	-10 3 5	20	-4 m	-4 3 7	22	-1 m	4 h 2	21	57 m	-3 h 5	29	-59 m
7 3 3	35	-48	-9 3 5	24	-4 m	-3 3 7	65	-36 m	5 h 2	85	-83 m	-2 h 5	60	-67 m
8 3 3	178	-174	-8 3 5	100	127 m	-2 3 7	22	-25 m	6 h 2	23	61 m	-1 h 5	60	-96 m
9 3 3	273	-282	-7 3 5	102	125 m	-1 3 7	22	-14 m	7 h 2	101	-102 m	0 h 5	60	-70 m
10 3 3	60	-32	-6 3 5	106	126 m	0 3 7	22	-17 m	8 h 2	136	-131 m	1 h 5	29	23 m
11 3 3	88	-59	-5 3 5	23	-19 m	1 3 7	22	-2 m	9 h 2	27	37 m	2 h 5	73	-91 m
12 3 3	24	-14 m	-4 3 5	39	49 m	2 3 7	22	15 m	10 h 2	129	125 m	3 h 5	30	57 m
13 3 3	98	-68	-3 3 5	23	22 m	3 3 7	22	22 m	11 h 2	29	-10 m	4 h 5	95	93 m
14 3 3	24	6 m	-2 3 5	23	-21 m	4 3 7	22	29 m				5 h 5	30	25 m
15 3 3	23	42 m	-1 3 5	23	-27 m	5 3 7	72	40 m						

A three-dimensional Fourier synthesis was calculated; a composite electron density map as viewed along the  $b$ -axis is shown in Fig. 1. A corresponding difference Fourier synthesis contained no electron density exceeding  $0.5 \text{ e}\cdot\text{\AA}^{-3}$ .

Final positional parameters and their standard deviations are listed in Table 1. The thermal parameters of the heavier atoms are listed in Table 2.

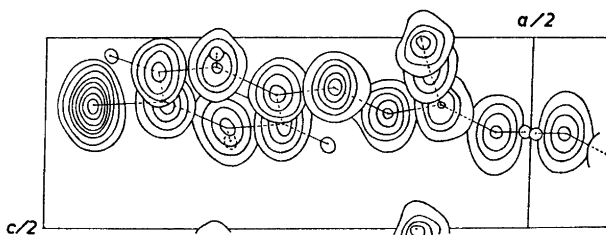


Fig. 1. Composite electron density map as viewed along the  $b$ -axis. Contour intervals are  $1.5 \text{ e}\cdot\text{\AA}^{-3}$ .

A comparison of observed and calculated structure factors is given in Table 3, the figures for non-observed reflections being those used in the least-squares calculations.

Table 4. Interatomic distances (Å) and bond angles (°). Standard deviations in parenthesis.

Cl—C(4)	1.735 (.007)	Cl—C(4)—C(3)	119.1 (.6)
C(1)—C(2)	1.381 (.011)	Cl—C(4)—C(5)	119.3 (.5)
C(2)—C(3)	1.368 (.010)	C(1)—C(2)—C(3)	119.5 (.5)
C(3)—C(4)	1.391 (.009)	C(2)—C(3)—C(4)	119.6 (.7)
C(4)—C(5)	1.374 (.012)	C(3)—C(4)—(5)	121.6 (.6)
C(5)—C(6)	1.395 (.010)	C(4)—C(5)—C(6)	119.4 (.6)
C(1)—C(6)	1.400 (.009)	C(5)—C(6)—C(1)	118.3 (.7)
C(1)—C(4)	2.734 (.009)	C(6)—C(1)—C(2)	121.7 (.6)
C(2)—C(5)	2.786 (.009)	N(1)—C(1)—C(2)	124.4 (.6)
C(3)—C(6)	2.787 (.012)	N(1)—C(1)—C(6)	113.9 (.7)
C(1)—N(1)	1.447 (.010)	C(1)—N(1)—N(2)	112.9 (.7)
N(1)—N(2)	1.264 (.010)	N(1)—N(2)—C(7)	111.2 (.8)
N(2)—C(7)	1.441 (.010)	N(2)—C(7)—N(4)	114.5 (.8)
C(7)—C(8)	1.432 (.013)	N(2)—C(7)—C(8)	119.1 (.6)
C(8)—N(3)	1.136 (.014)	N(4)—C(7)—C(8)	126.4 (.8)
C(7)—N(4)	1.273 (.009)	C(7)—C(8)—N(3)	177.2 (.8)
C(2)—H(1)	1.05 (.09)		
C(3)—H(2)	0.99 (.11)		
C(5)—H(3)	1.01 (.08)		
C(6)—H(4)	0.90 (.12)		
N(4)—H(5)	0.94 (.11)		

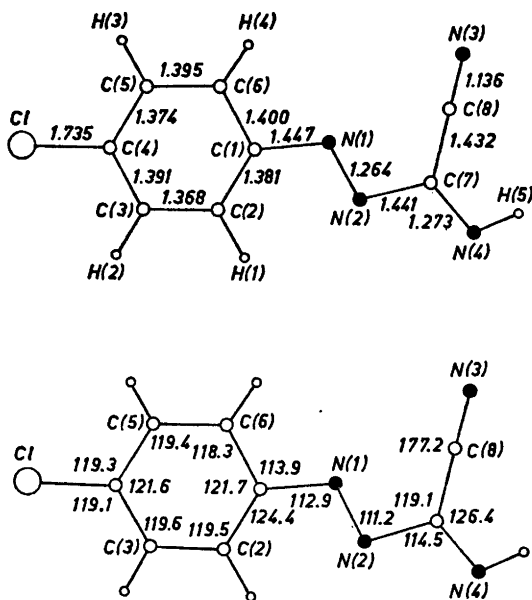


Fig. 2. Interatomic distances and bond angles.

## DISCUSSION

The structure determination demonstrates that Gabriel's "double cyanide" is actually an imidoglyoxynitrile of the aryldiazo radical. The bond distances and interbond angles are given in Table 4 together with their standard deviations. Distances and angles are also shown in Fig. 2.

The molecule is essentially planar. A least-squares plane containing the Cl, C, and N atoms has the equation  $-3.129x + 1.811y + 4.497z - 1.0 = 0$ , where  $x$ ,  $y$ , and  $z$  are fractional coordinate values. The nitrile nitrogen atom, N(3), is situated 0.08 Å from this plane; all other atoms are closer to it. An analysis of the equations of various planes through groups of atoms indicates a small twist ( $3.5^\circ$ ) about the C(1)—N(1) bond relative to a planar molecule, and an additional twist ( $2^\circ$ ) about the N(2)—C(7) bond.

Standard deviations of the  $B_{ij}$  values given in Table 2 are 5–10 % for  $i=j$  and up to 50 % for  $i \neq j$ . This comparatively large uncertainty and the fact that the intensity data have been collected mainly from one layer line set make a thermal vibrational analysis appear rather inconclusive. The mean of the atomic vibration seems, however, to increase markedly with the distance from the centre of gravity of the molecule, indicating a molecular libration about this point in addition to the lattice vibrations.

The geometry of the *p*-chlorobenzene-*anti*-diazo part of the molecule is in good agreement with the results from investigations of similar systems, e.g. *p*-chlorobenzene-*anti*-diazocyanide,<sup>7</sup> azobenzene,<sup>8</sup> and *p*-azotoluene.<sup>9</sup> A small deviation from hexagonal symmetry of the carbon arrangement in the benzene ring is indicated by an opening of the C(2)—C(1)—C(6) and C(3)—C(4)—C(5) angles, and by a significant shortening of the C(1)—C(4) diagonal (2.734 Å) relative to the two other diagonals (2.786 Å and 2.787 Å). The C—N distances in the diazo group are equal to 1.447 Å and 1.441 Å, as compared to 1.426 Å in the diazocyanide, 1.434 Å in azobenzene, and 1.433 Å in *p*-azotoluene; the N=N bond lengths is 1.264 Å, the corresponding values in the other compounds are 1.249 Å, 1.243 Å, and 1.244 Å, respectively. The non-linearity of C(4)—C(1)—N(1) is significant, and probably caused by repulsion between the N(2) and the C(2) and H(1) atoms.

The observed bond angles at the N(1) and N(2) atoms are  $112.9^\circ$  and  $111.2^\circ$ , respectively, which seem to be normal values for nitrogen bonded to *two* other atoms<sup>8-11</sup> with a double and a single bond.

The interbond angles at the C(7) atom may appear somewhat unexpected for an  $sp^2$  hybridized carbon atom. A certain deviation from the usually observed values might, however, be caused by intermolecular interactions. Bond distances in the imidoglyoxynitrile group are, the experimental uncertainty being taken into account, those to be expected.

Chains of hydrogen bonded molecules are present in the crystal; the imide nitrogen atom acts both as donor and acceptor. The N—N distance is 3.261 Å (e.s.d. 0.011 Å), and the hydrogen atom is slightly displaced (0.2 Å) from the line connecting the two nitrogen atoms.

A rather short distance between the C(2) atom in one molecule and the N(3) atom of a neighbouring molecule, 3.53 Å, is observed. The C(2)—(H(1))—



N(3)' direction is approximately normal to the nitrile triple bond (N(3)'—C(8)').

No other intermolecular distances are shorter than those corresponding to van der Waals' contacts.

## REFERENCES

1. Gabriel, S. *Ber.* **12** (1879) 1637.
2. Hantzsch, A. and Schulze, O. W. *Ber.* **28** (1895) 2073.
3. Hantzsch, A. and Danziger, K. *Ber.* **30** (1897) 2529.
4. Zollinger, H. *Chemie der Azofarbstoffe*, Birkhäuser Verlag, Basel 1958, p. 71.
5. Hamilton, W. C. *Acta Cryst.* **8** (1955) 185.
6. Hanson, H. P., Herman, F., Lea, J. D. and Skillman, S. *Acta Cryst.* **17** (1964) 1040.
7. Gram, F. and Rømming, Chr. *Selected Topics in Structure Chemistry*, Oslo 1967, p. 175.
8. Brown, C. J. *Acta Cryst.* **21** (1966) 146.
9. Brown, C. J. *Acta Cryst.* **21** (1966) 153.
10. Eichhorn, E. L. *Acta Cryst.* **12** (1959) 746.
11. Folting, K., Lipscomb, W. N. and Jerslev, B. *Acta Cryst.* **17** (1964) 1263.

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